



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CHEMISTRY

0620/62

Paper 6 Alternative to Practical

May/June 2016

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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[Turn over

Page 2	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – May/June 2016	0620	62

Abbreviations used in the Mark Scheme

- ; separates marking points
- / separates alternatives within a marking point
- **OR** gives alternative marking point
- **R** reject
- **I** ignore mark as if this material was not present
- **A** accept (a less than ideal answer which should be marked correct)
- **COND** indicates mark is conditional on previous marking point
- owtte or words to that effect (accept other ways of expressing the same idea)
- max indicates the maximum number of marks that can be awarded
- ecf credit a correct statement that follows a previous wrong response
- () the word / phrase in brackets is not required, but sets the context
- ora or reverse argument

Page 3	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – May/June 2016	0620	62

Question	Answer	Marks
1(a)	stand; beaker;	2 1 1
1(b)	arrow(s) underneath copper oxide;	1
1(c)	black; to orange / red / brown / pink;	2 1 1
1(d)	to condense (the water vapour);	1
1(e)(i)	water;	1
1(e)(ii)	test: anhydrous copper(II) sulfate; result: turns blue; OR test: cobalt(II) chloride (paper); result: turns pink;	2 1 1 1 1
1(e)(iii)	boiling / melting point determination;	1

Page 4	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – May/June 2016	0620	62

Question	Answer	Marks
2(a)	all 6 times completed correctly (2 marks) (22, 43, 64, 86, 105, 126) 5 times completed correctly (1 mark); in seconds;	3 2 1
2(b)	appropriate scale for y-axis / increasing at 20 s per large square; y-axis is a linear scale; all 6 points plotted correctly \pm half a small square (2 marks); 5 points plotted correctly \pm half a small square (1 marks); best-fit straight-line graph;	5 1 1 2 1
2(c)(i)	value from graph \pm half a small square (typically 167–170); units / s; extrapolation;	3 1 1 1
2(c)(ii)	sketch line below original line and diverging;	1
2(d)	as an indicator;	1
2(e)(i)	(more) accurate;	1
2(e)(ii)	solution slow to run out of pipette; difficult to know when to start timer / reaction does not start at once / inaccurate time measurement owtte;	2 1 1
2(f)	difficulty in swirling / mixing / shaking;	1

Page 5	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – May/June 2016	0620	62

Question	Answer	Marks
3(a)	<u>white</u> (solid / crystals / powder);	1
3(b)(i)	no change;	1
3(b)(ii)	turns from purple / pink; to colourless / white;	2 1 1
3(c)	yellow / orange (flame);	1
3(d)	ammonia / NH ₃ ;	1
3(e)	ammonium / NH ₄ ⁺ ;	1

Page 6	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – May/June 2016	0620	62

Question	Answer	Marks
4	<p>making the salt</p> <p>any 4 from:</p> <ul style="list-style-type: none"> • known volume sulfuric acid; • add named indicator; • add potassium hydroxide solution to the acid until the indicator changes colour/is neutralised; • note/measure the volume of potassium hydroxide solution added; • repeat without indicator OR add (decolourising) charcoal; <p>obtaining crystals</p> <p>any 2 from:</p> <ul style="list-style-type: none"> • heat/evaporate solution to crystallising point <u>until half evaporated</u> OR <u>until crystals (start to) form</u> OR <u>until saturated</u>; • leave to cool; • filter to get crystals; • dry crystals (on filter paper)/leave to dry; 	6